

Sodium

$^{23}_{11}\text{Na}$

Na^\bullet

$\frac{11}{11}$ p+

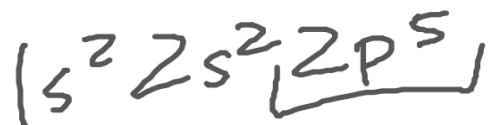
$\frac{12}{12}$ n^o

$\frac{11}{11}$ e-

$\frac{1}{1}$ ne-

$1s^2 2s^2 2p^6 3s^1$

.



Fluorine

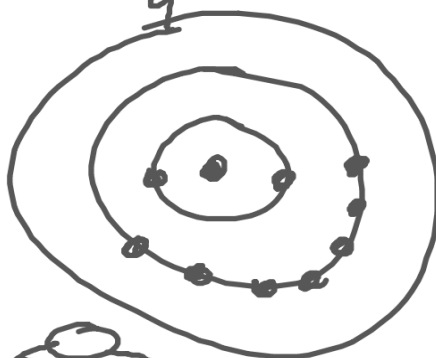
$^{19}_9\text{F}$

p⁺ 9

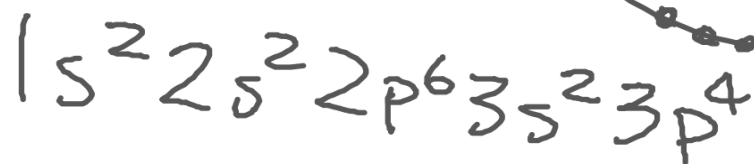
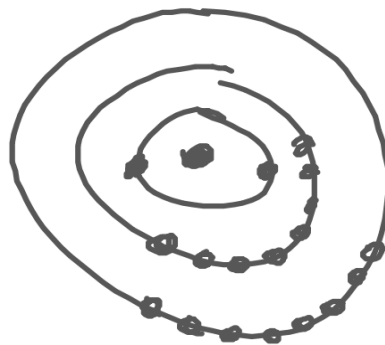
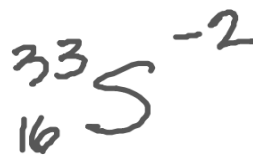
n^o 10

e⁻ 9

ne⁻ 7

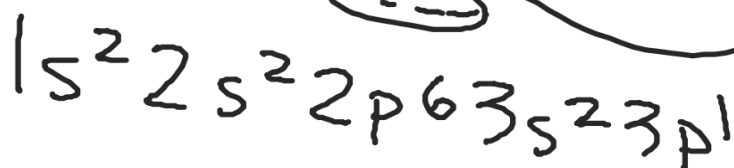
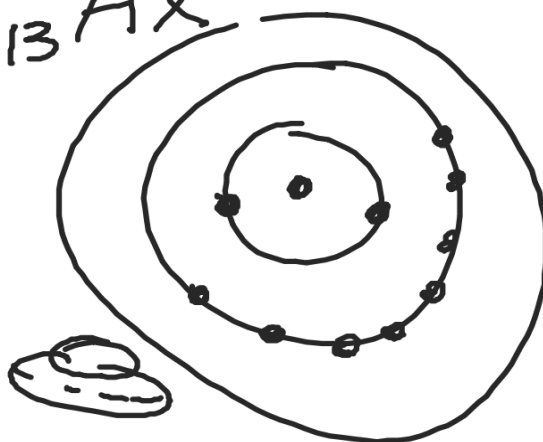


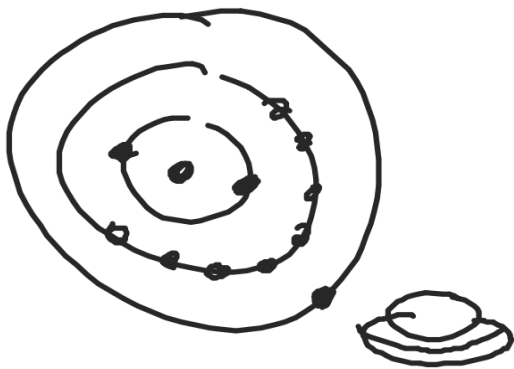
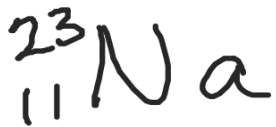
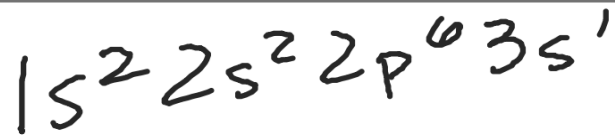
$$\begin{array}{r}
 p^+ \quad 16 \\
 \hline
 n^0 \quad 17 \\
 \hline
 e^- \quad 18 \\
 \hline
 Ne^- \quad 8
 \end{array}$$



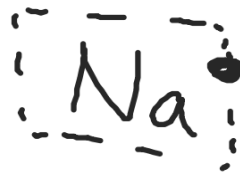
$$\begin{array}{r} 13 \text{ } p^+ \\ \hline 15 \text{ } n^0 \\ \hline 10 \text{ } e^- \\ \hline 8 \text{ } Ne^- \end{array}$$

$^{28}_{13}Al^{+3}$

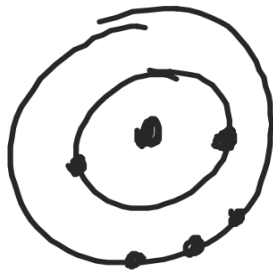




$$\begin{aligned} p^+ &= 11 \\ n^0 &= 12 \\ e^- &= 11 \\ \nu e^- &= 1 \end{aligned}$$







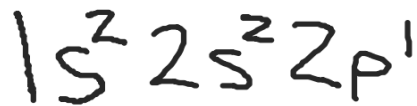
${}_{5}^{11}\text{B}$

$$p^{+} = 5$$

$$n^{0} = 6$$

$$e^{-} = 5$$

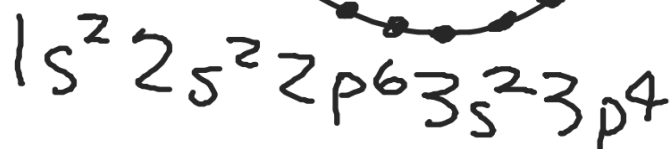
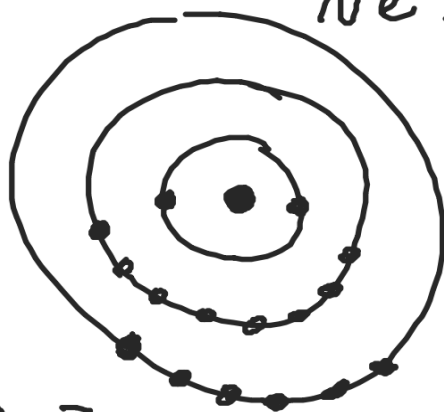
$$\underline{\underline{Ne^{-}}} = 3$$





$$\begin{array}{r} p+ \quad 16 \\ no \quad 16 \\ e- \quad 16 \\ ne- \quad 6 \end{array}$$

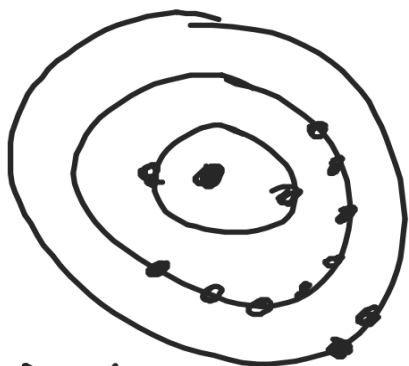
$^{32}_{16}S$



Mg[•]

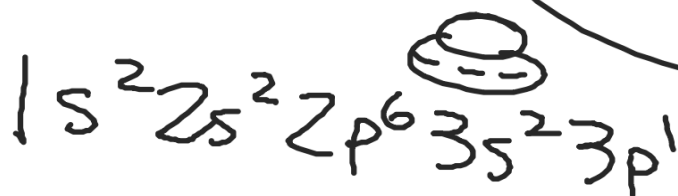
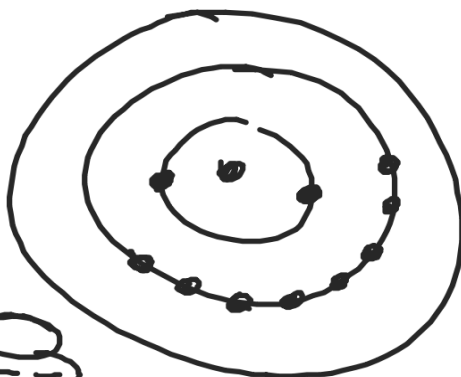
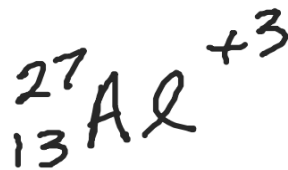
p⁺ 12
n^o 12
e⁻ 12
re⁻ 2

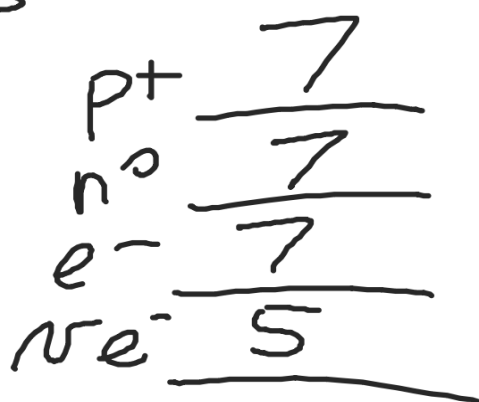
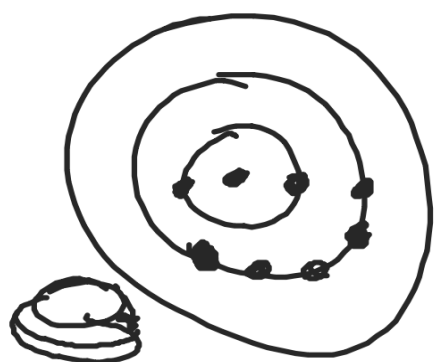
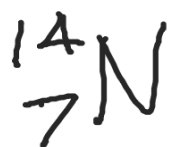
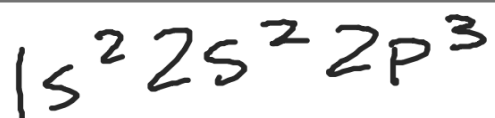
²⁴₁₂Mg

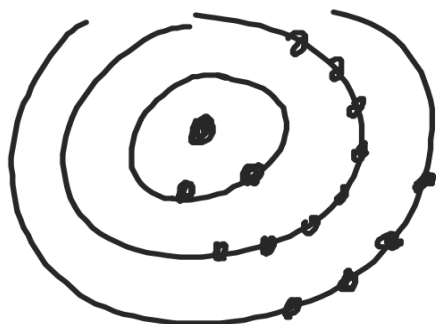


1s² 2s² 2p⁶ 3s²

$$\begin{array}{r}
 p^+ \quad 13 \\
 n^0 \quad 14 \\
 e^- \quad 10 \\
 \hline
 ve^- \quad 8 \\
 \hline
 \end{array}$$







$^{28}_{14}\text{Si}$

$p + \underline{14}$
 $n = \underline{14}$
 $e^- \underline{14}$
 $\text{Ne}^- \underline{4}$



$1s^2 2s^2 2p^6 3s^2 3p^2$



$$p^+ \frac{10}{}$$

$$n^0 \frac{10}{}$$

$$e^- \frac{10}{}$$

$$ne^- \underline{\underline{8}}$$

